



Cite this: *Phys. Chem. Chem. Phys.*,
2018, 20, 16167

The electronic structure and the nature of the chemical bond in CeO₂

Konstantin I. Maslakov,^a Yury A. Teterin,^{ab} Mikhail V. Ryzhkov,^{bc}
Aleksej J. Popel,^{cd} Anton Yu. Teterin,^b Kirill E. Ivanov,^b Stepan N. Kalmykov,^{ab}
Vladimir G. Petrov,^a Peter K. Petrov^e and Ian Farnan^d

The X-ray photoelectron spectral structure of CeO₂ valence electrons in the binding energy range of 0 to ~50 eV was analyzed. The core-electron spectral structure parameters and the results of relativistic discrete-variational calculations of CeO₈ and Ce₆₃O₂₁₆ clusters were taken into account. Comparison of the valence and the core-electron spectral structures showed that the formation of the inner (IVMO) and the outer (OVMO) valence molecular orbitals contributes to the spectral structure more than the many-body processes. The Ce 4f electrons were established to participate directly in chemical bond formation in CeO₂ losing partially their f character. They were found to be localized mostly within the outer valence band. The Ce 5p atomic orbitals were shown to participate in the formation of both the inner and the outer valence molecular orbitals (MOs). A large part in the IVMO formation is taken by the filled Ce 5p_{1/2}, 5p_{3/2} and O 2s atomic shells, while the Ce 5s electrons participate weakly in the chemical bond formation. The composition and the sequent order of the molecular orbitals in the binding energy range of 0 to ~50 eV were established. A quantitative scheme for the molecular orbitals of CeO₂ was built. This scheme is fundamental for understanding the nature of chemical bonding and also for the interpretation of other X-ray spectra of CeO₂. Evaluations revealed that the IVMO electrons weaken the chemical bond formed by the OVMO electrons by 37%.

Received 5th March 2018,
Accepted 25th April 2018

DOI: 10.1039/c8cp01442f

rsc.li/pccp

1 Introduction

Cerium dioxide is known as a non-radioactive structural substitute of actinide oxides (UO₂ and PuO₂).^{1,2} CeO₂-based ceramics are suggested as an inert ²³⁹Pu or ²³⁵U bearing matrix for nuclear fuel, as well as a matrix for high-level waste disposal.^{1,3} CeO₂ is also used as an exhaust gas afterburning catalyst^{4,5} and in electronics.⁶

X-ray photoelectron spectroscopy (XPS) has been widely used for the study of the electronic structure of CeO₂.^{7–11} The core and valence XPS spectra of cerium compounds are complicated and structured. This complicates the determination of the oxidation state of cerium on the basis of chemical shifts.^{7–11} Several mechanisms for the formation of the complex CeO₂ XPS structure were considered.

Thus, the authors of ref. 12–16 associated the complex CeO₂ XPS structure with the presence of Ce³⁺ and Ce⁴⁺ ions in the samples (mixed valence). However, the authors of ref. 17 suggested that CeO₂ contains only Ce⁴⁺ ions and on the basis of the calculations they associated the complex Ce 3d XPS structure with the complex final state consisting of the ground final state Ce 3d⁹4f⁰ and the two excited final states 3d⁹4f¹OVMO^{−1} and 3d⁹5p⁵np¹. The unusual final state Ce 3d⁹5p⁵np¹ was attributed to a weakly screened vacancy in the IVMO.^{9,18} These results agree with the experimental core-electron XPS data in ref. 19, which showed that the complex XPS structure appears in the core-electron XPS spectra and does not appear in the valence XPS spectra.

The authors of ref. 18 performed the non-relativistic discrete-variational calculations for the CeO₈ cluster including a cerium atom and its nearest environment in CeO₂ in the MO binding energy (BE) range of 0–22 eV and qualitatively compared the calculated and experimental spectra. In the present work the valence and the core-level XPS spectra of the single-crystalline CeO₂ film on Si(001) were measured, and the relativistic discrete-variational calculations for the CeO₈ and Ce₆₃O₂₁₆ clusters in the BE range of 0 to ~50 eV were carried out. A quantitative comparison of the calculated and experimental results was performed. A quantitative scheme of the molecular orbitals for CeO₂ was built.

^a Chemistry Department, Lomonosov Moscow State University, Moscow, 119991, Russia

^b NRC “Kurchatov Institute”, Moscow, 123182, Russia

^c Institute of Solid State Chemistry, Ural Department of RAS, Ekaterinburg, Russia

^d Department of Earth Sciences, University of Cambridge, Downing Street, Cambridge, CB2 3EQ, UK. E-mail: apopel@cantab.net; Tel: +44 (0)1223 768357

^e Department of Materials and London Centre for Nanotechnology, Imperial College London, London, SW7 2AZ, UK



2 Experimental

2.1 Sample production

The CeO₂ thin film was grown by pulsed laser deposition on a (001) oriented p-doped Si substrate and characterised by XRD and SEM as described in detail in ref. 20. The thin film of CeO₂ has nominal thickness (250 nm). X-ray diffraction analysis suggested that the CeO₂ film can be described as a single crystal in a (111) crystallographic orientation.

2.2 X-ray photoelectron measurements

The XPS spectra of CeO₂ were recorded on a Kratos Axis Ultra DLD spectrometer using monochromatic Al-K α radiation ($h\nu = 1486.7$ eV) at a 150 W X-ray gun power under 1.3×10^{-7} Pa at room temperature. The analyzed area was an ellipse with minor and major axes of 300 and 700 nm, respectively.

A Kratos charge neutraliser system was used and BEs were measured relatively to the BE of the C 1s electrons from hydrocarbons adsorbed on the sample surface, which was found to be equal to 285.0 eV. FWHMs are given relatively to that of the C 1s XPS peak for hydrocarbons on the sample surface to be 1.3 eV.²¹ The error in the determination of the BE and the peak width did not exceed ± 0.05 eV, and the error of the relative peak intensity was $\pm 5\%$. The inelastically scattered electron-related background was subtracted by the method of Shirley.²²

Quantitative elemental and ionic analysis was performed for the surface of the studied sample AP1 on the basis of the valence and core-level XPS spectra. The stoichiometric (CeO_{1.99}) and ionic (97% Ce⁴⁺ and 3% Ce³⁺) compositions of the studied sample were found. More details on the spectral measurements and elemental and ionic analyses are given in ref. 11.

2.3 Calculations

The calculations of the electronic structure of CeO₂ ($R_{\text{Ce-O}} = 0.2344$ nm)²³ were performed for the two types of finite fragments of the crystal lattice: 279-atom Ce₆₃O₂₁₆ and small CeO₈ clusters. The latter object modeled a cerium atom nearest environment in a dioxide crystal. For the modeling of the boundary conditions for the Ce₆₃O₂₁₆ fragment, we used an “extended cluster” scheme described in detail in ref. 24. In this model, the crystal fragment under study consists of two parts: the internal main part (or the “core”) and the outer part (or the “shell” of the cluster), the latter part usually includes the atoms of 1–5 coordination spheres surrounding the “core”. During the self-consistency procedure, the electron densities and the potential of atoms in the “shell” are replaced by the corresponding values obtained for the crystallographically equivalent centers of the cluster core. In addition to the introduction of the long-range component of the surrounding crystal potential, the extended cluster is embedded in a pseudopotential of the outer crystal lattice, including several thousand centers with Coulomb and exchange–correlation potentials obtained for the corresponding equivalent atoms in the internal part of the cluster. This pseudopotential modeling of the long-range crystal potential becomes important in the calculations of

compounds with noticeable charges on atoms, such as oxides and fluorides.

In the case of Ce₆₃O₂₁₆, the core of the cluster included: a Ce atom in the center (Ce₁) with the eight nearest oxygen neighbors (O₁) and 12 cerium sites of the next metal coordination sphere with their 48 nearest oxygen atoms. The other atoms of the cluster formed the “shell” and during self-consistency their electron densities and potentials were kept equivalent to those of Ce₁ and O₁. In the case of small fragment CeO₈, the renormalization procedure for the populations of oxygen atoms in valence atomic orbitals (AOs) during self-consistency was used in the calculations. The latter model of the boundary conditions for small clusters also allows one to include into the iterative scheme the stoichiometry of compounds and the possibility of charge redistribution between the outer atoms in the cluster and the surrounding crystal.

In this work, the calculations were carried out using the fully relativistic discrete-variational method (RDV)^{25,26} with local exchange–correlation potential.²⁷ The RDV method is based on the solution of the Dirac–Slater equation for 4-component wave functions transforming according to the irreducible representations of the double point group (D_{4h} in the present calculations). For the calculation of symmetry coefficients we used the original code which realizes the projection operator technique²⁵ and includes the matrices of the irreducible representations of double point groups obtained in ref. 28 and the transformation matrices presented in ref. 29. The extended basis of 4-component numerical atomic orbitals (AOs) obtained as the solution of the Dirac–Slater equation for the isolated neutral atoms also included the vacant Ce 6p_{1/2} and 6p_{3/2} functions in addition to completely or partially occupied AOs. Numerical Diophantine integration in the calculations of matrix elements was carried out for 698 000 (Ce₆₃O₂₁₆) and 22 000 (CeO₈) sample points distributed in the cluster space. This provided the convergence of MO energies within 0.1 eV. The calculations of the electronic structure have been proved to be one of the most accurate tools for the determination of the electronic structures of solids.^{30–32}

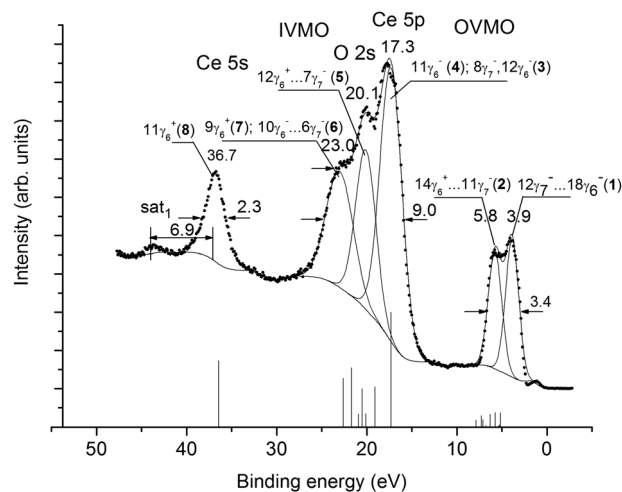
3 Results and discussion

The valence XPS structure of CeO₂ (0 to ~50 eV BE) appears due to OVMO and IVMO formation, as well as due to other processes (Fig. 1). In order to evaluate the contributions of the other processes, the core-level XPS structure was analyzed. For more details see ref. 11. The present work considers only some of the core-level XPS spectra.

3.1 Core-level XPS structure of CeO₂

As shown in ref. 11, the Ce 3p_{1/2}, 3d-, 4s-, 4p-, 4d-, 5s spectra in the BE range from ~35 to ~1250 eV exhibit a complex structure due to the many-body perturbation and complex final states. After Ce 3d electron photoemission the ground final state Ce 3d⁹4f⁰ and the two excited final states 3d⁹4f¹OVMO⁻¹ and 3d⁹5p⁵np¹ appear. The final state Ce 3d⁹5p⁵np¹ was





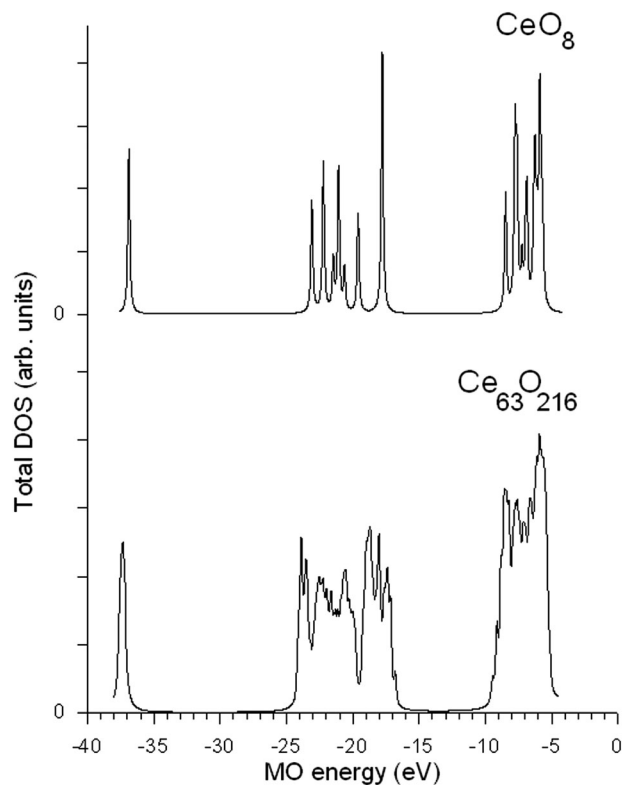


Fig. 4 Total DOS for the central part of CeO_8 and $\text{Ce}_{63}\text{O}_{216}$ clusters.

for the illustration of the general structure of the main molecular states, we use the results of RDV calculations for the small CeO_8 cluster. It is worth mentioning that in Fig. 1, the theoretical MO intensities were calculated using photoionization cross-sections.^{33,34} The characteristics of all vacant and occupied MOs in the energy region from -40 to $+15$ eV obtained for this cluster are given in Table 1. The MO LCAO (Molecular Orbitals as Linear Combinations of Atomic Orbitals) approach was employed in order to consider the chemical bond in terms of atomic and molecular shells.

The IVMO and the OVMO formation is a result of the hybridization of Ce and O orbitals (Ce 5s, 5p, 4f, 6s and O 2s, 2p and empty Ce 5d, 6p AOs). Unlike the results of non-relativistic calculations for CeO_2 showing a significant participation of Ce 5s AO in MO formation,¹⁸ the relativistic calculations show that Ce 5s states do not hybridize with oxygen orbitals (Table 1). The Ce 6s and 6p AOs take a notable part in the formation of vacant MOs. The Ce 4f AOs mostly participate in OVMO formation, while the Ce 5p, 5d AOs participate in both the OVMO and the IVMO. Especially a high mixing of the Ce $5p_{3/2}$ and the O 2s AOs was observed for the $12\gamma_6^-$, $8\gamma_7^-$ (3) and $10\gamma_6^-$, $6\gamma_7^-$ (6) IVMOs (Table 1). The hybridization of the Ce $5p_{1/2}$ and O 2s AOs in the $11\gamma_6^-$ (4) and $9\gamma_6^-$ (7) IVMOs is even higher due to a decrease in the BE difference between Ce $5p_{1/2}$ and O 2s levels (Fig. 1). These results together with the photoionization cross-sections^{33,34} allow one to build a theoretical XPS spectrum (Table 2) and to interpret the corresponding experimental valence XPS spectral structure of CeO_2 (Fig. 1).

3.3 Valence XPS spectrum of CeO_2

The valence XPS spectrum exhibits several explicit features. This BE range can be subdivided into components (Fig. 1). Although formalistic, this subdivision allows one to compare the theoretical and experimental spectra of CeO_2 (Fig. 1 and Table 2).

The experimental spectrum reflects the excited electronic state with a hole in a level, while the theoretical spectrum corresponds to the ground state. Therefore, the BE values calculated for the transition state should be used.³⁵ However, it is known that the valence BEs calculated for the transition state usually differ from those for the ground state by a constant shift toward a higher value. Therefore, the BEs (Table 1) are increased by 5.14 eV (Table 2). With the MO composition (Table 1) and photoionization cross-sections in mind,^{33,34} the intensities of the spectral ranges were calculated (Table 2). It has to be taken into account that CeO_2 XPS reflects the band structure and consists of bands widened due to the solid-state effects. Despite the approximation imperfection, a satisfactory quantitative agreement between the theoretical and experimental data was observed (Fig. 1 and Table 2).

The OVMO band can be subdivided into two components (1, 2) at 3.9 and 5.8 eV (Fig. 1 and Table 2). Component (1) is mostly due to the Ce 4f and O 2p electrons, while component (2) is due to the Ce 5d and O 2p electrons (Table 1). This agrees with the results of non-relativistic calculations for the CeO_8 cluster¹⁸ and with the resonant emission (OVMO \rightarrow Ce 4d⁹) spectroscopy data³⁶ for CeO_2 . Under the resonant excitation at 124.8 eV, component (1) grows considerably at 3.56 eV. This confirms that this peak reflects the Ce 4f electronic states. The resonant emission spectrum of CeO_2 does not exhibit a feature at a lower energy side from the OVMO band associated with the Ce 4f electrons, whereas the corresponding spectrum of Ce_2O_3 exhibits this feature at 1.95 eV. This indicates that a stoichiometric CeO_2 consists of only Ce^{4+} ions. Thus, the Ce 4f related feature at 1.1 eV in the spectrum of CeO_2 (Fig. 1) appears due to Ce^{3+} ions formed on the surface under the Al K_α X-ray beam. In previous papers, the increase in the Ce^{3+} concentration on the surface of CeO_2 under the Al K_α X-ray beam^{36,37} and increased temperature^{5,38,39} were observed.

The IVMO spectrum can be formally subdivided into several ranges (Fig. 1 and Table 2). The Ce 5s spectrum does not show a 40% high-BE satellite with $\Delta E_{\text{sat}2} = 15$ eV, the intensity of which was taken into account for the determination of the experimental intensity (15%) of the $11\gamma_6^+$ IVMO (Table 2). The calculated ($I_{\text{T}}(\text{OVMO}) = 2.72$ eV, $I_{\text{T}}(\text{IVMO}) = 18\%$ and $I_{\text{T}}(\text{IVMO}) = 82\%$) and the experimental ($I_{\text{E}}(\text{OVMO}) = 3.4$ eV, $I_{\text{E}}(\text{OVMO}) = 22\%$ and $I_{\text{E}}(\text{IVMO}) = 79\%$) data are comparable (Table 2 and Fig. 1). A satisfactory agreement between the theoretical and experimental BEs of some electronic shells was reached (Table 2). As noted, the non-relativistic calculations¹⁸ allowed only a qualitative interpretation of the low-BE part of the valence XPS spectrum of CeO_2 . The results of the present relativistic calculations allow a



Table 1 MO composition, energy E_0^a (eV) and the photoionization cross-sections σ_i^b for the CeO₈ cluster

		MO composition												
		Ce										O		
		5s	5p _{1/2}	5p _{3/2}	5d _{3/2}	5d _{5/2}	6s	4f _{5/2}	4f _{7/2}	6p _{1/2}	6p _{3/2}	2s	2p _{1/2}	2p _{3/2}
MO	$-E_0$ (eV)	σ_i^a 1.56	1.51	1.50	0.47	0.44	0.15	0.97	0.93	0.89	0.86	0.96	0.07	0.07
OVMO	19 γ_7^-	-14.97		0.01							0.9	0.04	0.02	0.03
	23 γ_6^-	-14.95		0.01						0.01	0.9	0.04	0.01	0.03
	22 γ_6^-	-14.79	0.01							0.89	0.01	0.04	0.02	0.03
	22 γ_6^+	-12.63					0.88					0.05	0.02	0.05
	19 γ_7^+	-11.77				0.81						0.04	0.09	0.06
	18 γ_7^+	-11.59			0.45	0.36						0.04	0.04	0.11
	21 γ_6^+	-11.58			0.46	0.35						0.04	0.02	0.13
	20 γ_6^+	-7.5			0.38	0.50							0.04	0.08
	17 γ_7^+	-7.5			0.38	0.50							0.04	0.08
	18 γ_7^-	-4.17						0.18	0.69			0.01	0.05	0.07
	21 γ_6^-	-3.71							0.94	0.01			0.05	
	17 γ_7^-	-3.68						0.01	0.94				0.01	0.04
	20 γ_6^-	-3.68						0.01	0.94				0.01	0.04
	16 γ_7^-	-3.41						0.74	0.22				0.01	0.03
	19 γ_6^-	-3.35						0.92	0.02		0.01			0.05
	15 γ_7^-	-3.35						0.93	0.02					0.05
	18 γ_6^- ^c	0											0.31	0.69
	14 γ_7^-	0											0.31	0.69
	13 γ_7^-	0.13		0.04				0.04	0.02		0.01		0.21	0.68
	17 γ_6^-	0.13		0.04				0.04	0.01		0.01		0.13	0.77
	19 γ_6^+	0.22											0.16	0.84
	16 γ_7^+	0.22											0.15	0.85
	18 γ_6^+	0.23											0.66	0.34
	16 γ_6^-	0.58	0.02						0.05				0.72	0.21
	15 γ_7^+	0.58				0.01							0.06	0.93
	14 γ_7^+	0.6			0.01	0.01							0.48	0.50
	17 γ_6^+	0.6			0.01	0.01							0.46	0.52
	12 γ_7^-	0.66						0.06	0.07				0.29	0.58
	11 γ_7^-	1.14		0.01				0.01	0.01		0.04		0.25	0.68
	15 γ_6^-	1.14		0.01				0.01			0.04		0.13	0.81
	14 γ_6^+	1.27		0.01						0.05			0.44	0.50
	16 γ_6^+	1.65	0.01				0.06					0.01	0.31	0.61
	10 γ_7^-	1.94						0.01	0.02				0.17	0.80
	13 γ_6^-	1.94						0.01	0.02				0.17	0.80
	9 γ_7^-	1.94						0.02	0.02				0.64	0.32
	13 γ_7^+	2.12			0.05	0.08						0.01	0.08	0.78
	15 γ_6^+	2.12			0.08	0.05						0.01	0.03	0.83
	12 γ_7^+	2.13			0.03	0.10						0.01	0.76	0.10
	11 γ_7^+	2.71			0.05	0.07							0.29	0.59
	14 γ_6^+	2.72			0.05	0.07							0.30	0.58
IVMO	12 γ_6^-	12.16		0.69							0.02	0.24	0.01	0.04
	8 γ_7^-	12.16		0.69							0.02	0.24	0.02	0.03
	11 γ_6^-	13.96		0.41						0.03		0.54	0.01	0.01
	7 γ_7^-	14.95						0.01	0.01			0.98		
	10 γ_7^+	15.37			0.01	0.03						0.96		
	9 γ_7^+	15.37			0.02	0.03						0.95		
	13 γ_6^+	15.37			0.03	0.02						0.95		
	12 γ_6^+	15.77	0.01				0.06					0.93		
	6 γ_7^-	16.55		0.25							0.02	0.71	0.01	0.01
	10 γ_6^-	16.55		0.25							0.02	0.72		0.01
	9 γ_6^-	17.45		0.55						0.01		0.42	0.01	0.01
	11 γ_6^+	31.31	0.98									0.01		0.01

^a Levels shifted by 13.01 eV toward the positive values (upward). ^b Photoionization cross-sections σ_i (kiloBarn per electron) from ref. 33. ^c Highest occupied MO (HOMO; 2 electrons), the occupation number for all the orbitals is 2.

satisfactory quantitative identification of the CeO₂ valence XPS spectrum in the BE range from 0 to ~50 eV (Table 2 and Fig. 1). These data will promote the development of the RDV method and help to understand the chemical bond nature and to construct a quantitative MO scheme in CeO₂.

3.4 Molecular orbital scheme of CeO₂

With the BE differences between the valence- and the core-level peaks, as well as the calculation results for atomic³⁴ Ce and the CeO₈ cluster, a quantitative scheme of the molecular orbitals for CeO₂ was built (Fig. 5). In the accepted approximation one



Table 2 Parameters of CeO₂ XPS spectra, CeO₈ cluster and Ce 5p and Ce 4f DOS ρ_i (e⁻)

MO		$-E^a$ (eV)	XPS			Ce 5p,4f DOS ρ_i e ⁻ (electron)			
			Energy ^b (eV)		Intensity, %	4f _{5/2}	4f _{7/2}	5p _{1/2}	5p _{3/2}
			Experiment	Theory					
OVMO	18 γ_6^- ^c	5.14	3.9 (1.7)	0.2	12				
	14 γ_7^-	5.14		0.2					
	13 γ_7^-	5.27		1.8		0.08	0.04		0.08
	17 γ_6^-	5.27		1.7		0.08	0.02		0.08
	19 γ_6^+	5.34		0.2					
	16 γ_7^+	5.34		0.2					
	18 γ_6^+	5.37		0.2					
	16 γ_6^-	5.72		0.9			0.1	0.08	
	15 γ_7^+	5.72		0.3					
	14 γ_7^+	5.74		0.3					
	17 γ_6^-	5.74		0.3					
	12 γ_7^-	5.80	5.8 (1.8)	1.8	10	0.12	0.14		
	11 γ_7^-	6.28		1.1		0.02	0.02		0.02
	15 γ_6^-	6.28		1.0		0.02			0.02
	14 γ_6^+	6.41		1.0				0.04	
	16 γ_6^+	6.79		0.6					
	10 γ_7^-	7.08		0.6		0.02	0.04		
	13 γ_6^-	7.08		0.6		0.02	0.04		
	9 γ_7^-	7.08		0.7		0.04	0.04		
	13 γ_7^+	7.26		1.0					
	15 γ_6^+	7.26		1.0					
	12 γ_7^+	7.27		0.9					
	11 γ_7^+	7.85		0.9					
	14 γ_6^+	7.86		0.9					
	ΣI_i^d			18.4	22	0.4	0.44	0.12	0.2
IVMO	12 γ_6^-	17.30	17.3 (2.9)	14.3	38				1.38
	8 γ_7^-	17.30		14.3					1.38
	11 γ_6^-	19.10		10.0				0.82	
	7 γ_7^-	20.09	20.1 (2.6)	3.3	20	0.02	0.02		
	10 γ_7^+	20.51		3.2					
	9 γ_7^+	20.51		3.2					
	13 γ_6^+	20.51	23.0 (3.5)	3.2	20				
	12 γ_6^+	20.91		3.2					
	6 γ_7^-	21.69		7.4					0.5
	10 γ_6^-	21.69		7.4					0.5
	9 γ_6^-	22.59		12.1				1.10	
	ΣI_i^d			81.6	79	0.02	0.02	1.92	3.76
	11 γ_6^+	36.45	36.7 (2.4)	~16.5	15				

^a Energies shifted toward higher values (down) by 5.14 eV. ^b FWHMs are given in the parentheses in eV. ^c The highest occupied MO (2 electrons), the occupation number for other orbitals is 2. ^d Total intensities and Ce 5p, 4f DOS.

can formally separate the antibonding 12 γ_6^- , 8 γ_7^- (3) and 11 γ_6^- (4) and the corresponding bonding 10 γ_6^- , 6 γ_7 (6) and 9 γ_6^- (7) IVMOs, as well as the quasiatomic 7 γ_7^- , 10 γ_7^+ , 9 γ_7^+ , 13 γ_6^+ and 12 γ_6^+ (5) IVMOs associated mostly with the O 2s states which have to be close in magnitude.

Indeed, the O 1s spectrum of CeO₂ allows one to evaluate that the inequality of O 2s states must not exceed 1 eV, since the O 1s peak was observed to be sharp ($\Gamma = 1.0$ eV) and symmetric (Fig. 3). The O 2s BE must be $E_b(\text{O } 2s) = 21.1$ eV, since for atomic oxygen $\Delta E_O = E_b(\text{O } 1s) - E_b(\text{O } 2s) = 508.3$ eV³⁴ and the O 1s BE is $E_b = 529.4$ eV (Fig. 3). These data partially agree with the results of theoretical calculations (Table 2). Taking into account that the calculated BE difference for Ce atoms is $\Delta E_{\text{Ce}}^T = 864.5$ eV ($E_b(\text{Ce } 3d_{5/2}) = 883.3$ eV and $E_b(\text{Ce } 5p_{3/2}) = 18.8$ eV),³⁴ while the difference between the Ce 3d_{5/2} and 12 γ_6^- , 8 γ_7^- (3) IVMO is $\Delta E_1 = 865.0$ eV, one can find that $\Delta_1 = 0.5$ eV (Fig. 5).

This confirms the antibonding nature of the 12 γ_6^- , 8 γ_7^- (3) IVMO. The FWHM of the IVMO XPS features does not allow to arrive at a conclusion on their bonding or antibonding nature. However, one can suggest that these orbitals partially lose their antibonding nature (Table 2 and Fig. 5, see also ref. 40) due to the contribution of 5% of the O 2p and 2% of the Ce 6p AO in the 12 γ_6^- , 8 γ_7^- (3) IVMO.

3.5 Contribution of valence electrons to the chemical bond in CeO₂

In this work, we used the values of overlap populations (n_{ij}) of various pairs of metal and oxygen AOs⁴¹ to evaluate their contributions to the chemical bond. Table 3 shows these values for the outer and inner valence electrons obtained in the non-relativistic (NDV) and relativistic (RDV) calculations. Positive and negative values correspond to the bonding and



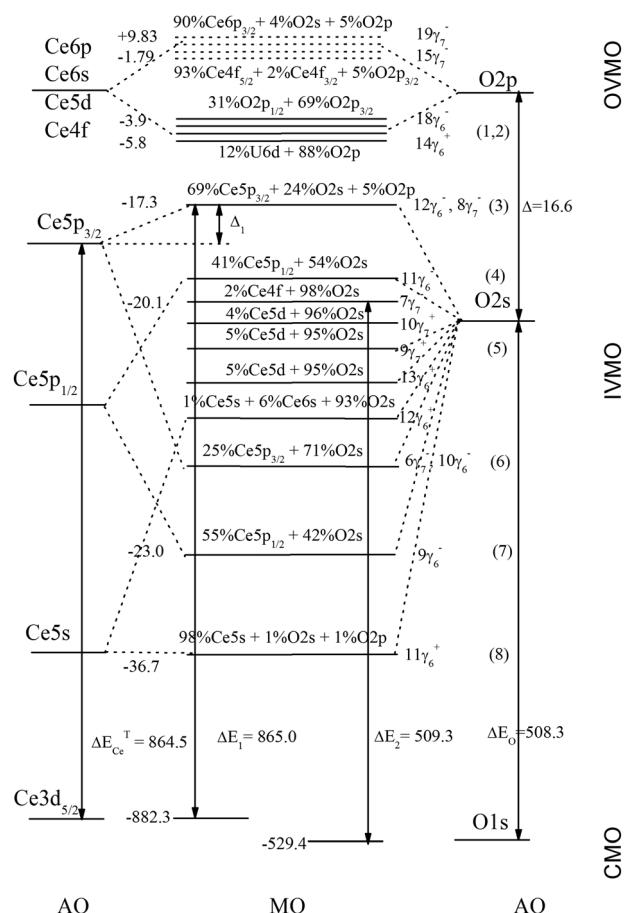


Fig. 5 Molecular orbital scheme for the CeO_8 cluster built on the basis of theoretical and experimental data. Chemical shift is not shown. Arrows indicate experimentally measured BE differences. Experimental BEs (eV) are shown on the left. The energy scale is not shown.

Table 3 Overlap populations for CeO_2 (per one ligand, $\times 10^3$), obtained in the relativistic (RDV) and non-relativistic (NDV) calculations of the CeO_8 cluster

Bonds in CeO_2	RDV ^a		NDV
	Partial	Total	
Ce $4f_{5/2}$ -O $2p_{1/2}$; O $2p_{3/2}$	1; 21	49	43
Ce $4f_{7/2}$ -O $2p_{1/2}$; O $2p_{3/2}$	15; 12		
Ce $4f_{5/2}$ -O 2s	3		
Ce $4f_{7/2}$ -O 2s	4	7	5
Ce $6p_{1/2}$ -O $2p_{1/2}$; O $2p_{3/2}$	18; 5		
Ce $6p_{3/2}$ -O $2p_{1/2}$; O $2p_{3/2}$	3; 42	68	67
Ce $6p_{1/2}$ -O 2s	11		
Ce $6p_{3/2}$ -O 2s	19	30	27
Ce $6s$ -O $2p_{1/2}$; O $2p_{3/2}$	4; 7	11	5
Ce $6s$ -O 2s	26	26	27
Ce $5d_{3/2}$ -O $2p_{1/2}$; O $2p_{3/2}$	4; 72		
Ce $5d_{5/2}$ -O $2p_{1/2}$; O $2p_{3/2}$	58; 55	189	179
Ce $5d_{3/2}$ -O 2s	12		
Ce $5d_{5/2}$ -O 2s	19	31	30
Σ_{OVMO}^b	411	411	383
Ce $5p_{1/2}$ -O $2p_{1/2}$; O $2p_{3/2}$	-8; -18		
Ce $5p_{3/2}$ -O $2p_{1/2}$; O $2p_{3/2}$	-23; -40	-89	-91
Ce $5p_{1/2}$ -O 2s	-7		
Ce $5p_{3/2}$ -O 2s	-21	-28	-31
Ce $5s$ -O $2p_{1/2}$; O $2p_{3/2}$	-10; -20	-30	-37
Ce $5s$ -O 2s	-3	-3	-5
Σ_{IVMO}^b	-150	-150	-164

^a Partial and total contribution. ^b Total OVMO and IVMO contribution.

antibonding interactions, respectively. The relativistic and non-relativistic results differ slightly. For example, the contribution from the Ce 4f-O 2p, Ce 6s-O 2p and Ce 5d-O 2p bonds increases in the relativistic approach, which reveals the bonding nature of these states. The Ce 5s-O 2s, 2p interactions are less antibonding in the non-relativistic calculation than in the relativistic one. This can be explained by the fact that in the relativistic calculation, the Ce 5s BE increases significantly compared to that calculated in the non-relativistic approach.

The contribution to bonding from the outer valence shells in CeO_2 is (411) (Table 3). The largest bonding contributions come from the Ce 4f-O 2p (49), Ce 6p-O 2p (68) and Ce 5d-O 2p (189) AOs. The inner valence electrons of Ce provide the total antibonding contribution of -150. The highest antibonding contribution is obtained from the Ce 5p-O 2p (-89) states. To summarize, the IVMO electrons weaken the chemical bond by 37%. The total contribution of the valence electrons to the chemical bond in CeO_2 is 261.

4 Conclusions

On the basis of the relativistic cluster calculations, a quantitative interpretation of the outer (0 to ~ 15 eV BE) and inner (~ 15 to ~ 50 eV BE) valence XPS spectral structures of CeO_2 was performed.

It was theoretically shown and experimentally confirmed that the Ce 4f electrons (0.86 of the Ce 4f electron) participate directly in chemical bond formation in CeO_2 partially losing the f-character. The contributions from Ce 4f orbitals are located mainly within the outer valence band.

It was found that the Ce 5p electrons not only participate effectively (experimentally measurable) in the formation of the IVMO, but also participate noticeably in the formation of the occupied OVMO (0.32 of the Ce 5p electron). The largest part in the IVMO formation is taken by the Ce $5p_{1/2}$, $5p_{3/2}$ and the O 2s AOs of the neighboring cerium and oxygen. The Ce 5s electrons participate weakly in the chemical bond formation.

The composition and the sequent order of the molecular orbitals in the binding energy range from 0 to ~ 50 eV were established. A quantitative scheme of the molecular orbitals for CeO_2 was built. This scheme is fundamental for understanding the nature of chemical bonding and for the interpretation of other X-ray spectra of CeO_2 . The evaluation of the contribution from various electrons to the chemical bond revealed that the antibonding contribution from IVMO electrons is near 37% of the chemical bond formed by the OVMO electrons.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The work was supported by the RFBR grant no. 17-03-00277a. The authors acknowledge support from Lomonosov Moscow



State University Program of Development for providing access to the XPS facility. M. V. Ryzhkov acknowledges financial support of FASO of Russia ISSC of the Ural Branch of RAS No. AAAA-A16-116122810214-9. A. J. Popel acknowledges funding from the UK EPSRC (grant EP/I036400/1) and Radioactive Waste Management Ltd (formerly the Radioactive Waste Management Directorate of the UK Nuclear Decommissioning Authority, contract NPO004411A-EPS02), a maintenance grant from the Russian Foundation for Basic Research (projects 13-03-90916) and CSAR bursary.

References

- 1 P. D. Edmondson, Y. Zhang, S. Moll, F. Namavar and W. J. Weber, Irradiation effects on microstructure change in nanocrystalline ceria – Phase, lattice stress, grain size and boundaries, *Acta Mater.*, 2012, **60**, 5408–5416.
- 2 M. Kinoshita, K. Yasunaga, T. Sonoda, A. Iwase, N. Ishikawa, M. Sataka, K. Yasuda, S. Matsumura, H. Y. Geng, T. Ichinomiya, Y. Chen, Y. Kaneta, M. Iwasawa, T. Ohnuma, Y. Nishiura, J. Nakamura and H. Matzke, Recovery and restructuring induced by fission energy ions in high burnup nuclear fuel, *Nucl. Instrum. Methods Phys. Res., Sect. B*, 2009, **267**, 960–963.
- 3 B. E. Burakov, M. I. Ojovan and W. E. Lee, Crystalline Materials for Actinide Immobilisation, in *Materials for Engineering*, ed. W. E. Lee, Imperial College Press, London, 2011, vol. 1, pp. 58–59.
- 4 D. R. Mullins, S. H. Overbury and D. R. Huntley, Electron spectroscopy of single crystal and polycrystalline cerium oxide surfaces, *Surf. Sci.*, 1998, **409**, 307–319.
- 5 V. Stetsovych, F. Pagliuca, F. Dvorak, T. Duchon, M. Vorokhta, M. Aulicka, J. Lachnitt, S. Schernich, I. Matolinova, K. Veltruska, T. Skala, D. Mazur, J. Myslivecek, J. Libuda and V. Matolin, Epitaxial Cubic Ce₂O₃ Films via Ce-CeO₂ Interfacial Reaction, *J. Phys. Chem. Lett.*, 2013, **4**, 866–871.
- 6 Y. Zhu, N. Jain, M. K. Hudait, D. Maurya, R. Varghese and S. Priya, X-ray photoelectron spectroscopy analysis and band offset determination of CeO₂ deposited on epitaxial (100), (110), and (111)Ge, *J. Vac. Sci. Technol., B*, 2014, **32**, 011217.
- 7 P. Burroughs, A. Hamnett, A. F. Orchard and G. Thornton, Satellite structure in the X-ray photoelectron spectra of some binary and mixed oxides of lanthanum and cerium, *J. Chem. Soc., Dalton Trans.*, 1976, 1686–1698.
- 8 Yu. A. Teterin and A. S. Baev, *Rentgenovskaya Fotoelektronnaya Spektroskopiya Soedinenii Lantanoidov (X-ray Photoelectron Spectroscopy of Lanthanide Compounds)*, TsNII atominform, Moscow, 1987 (in Russian).
- 9 Yu. A. Teterin and A. Yu. Teterin, Structure of X-ray photoelectron spectra of lanthanide compound, *Russ. Chem. Rev.*, 2002, **71**, 347–381.
- 10 Yu. A. Teterin, S. V. Stefanovskii, S. V. Yudinsev, G. N. Bek-Uzarov, A. Yu. Teterin, K. I. Maslakov and I. O. Utkin, X-ray photoelectron study of calcium cerium titanate ceramics, *Russ. J. Inorg. Chem.*, 2004, **49**, 87–94.
- 11 K. I. Maslakov, Yu. A. Teterin, A. J. Popel, A. Yu. Teterin, K. E. Ivanov, S. N. Kalmykov, V. G. Petrov, P. K. Petrov and I. Farnan, XPS study of ion irradiated and unirradiated CeO₂ bulk and thin film samples, *Appl. Surf. Sci.*, 2018, **448**, 154–162.
- 12 A. Kotani, T. Jo and J. C. Parlebas, Many-body effects in core-level spectroscopy of rare-earth compounds, *Adv. Phys.*, 1988, **37**, 37–85.
- 13 A. Kotani and H. Ogasawara, Theory of core-level spectroscopy of rare-earth oxides, *J. Electron Spectrosc. Relat. Phenom.*, 1992, **60**, 257–299.
- 14 A. Fujimori, Mixed-valent ground state of CeO₂, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1983, **28**, 2281–2283.
- 15 A. Kotani, Theory of core-level spectroscopy in f and d electron systems, *J. Electron Spectrosc. Relat. Phenom.*, 1999, **100**, 75–104.
- 16 E. Shoko, M. F. Smith and R. H. McKenzie, Mixed valency in cerium oxide crystallographic phases: Valence of different cerium sites by the bond valence method, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2009, **79**, 134108.
- 17 G. Thornton and M. J. Dempsey, Final-state effects in the 3d and 4d X-ray photoelectron spectra of CeO₂, *Chem. Phys. Lett.*, 1981, **77**, 409–412.
- 18 M. V. Ryzhkov, V. A. Gubanov, Yu. A. Teterin and A. S. Baev, Electronic structure, chemical bonding and X-ray photoelectron spectra of light rare-earth oxides, *Z. Phys. B: Condens. Matter*, 1985, **59**, 1–6.
- 19 Yu. A. Teterin, A. Yu. Teterin, A. M. Lebedev and I. O. Utkin, The XPS spectra of cerium compounds containing oxygen, *J. Electron Spectrosc. Relat. Phenom.*, 1998, **88–91**, 275–279.
- 20 A. J. Popel, S. Le Sollic, G. I. Lampronti, J. Day, P. K. Petrov and I. Farnan, The effect of fission-energy Xe ion irradiation on the structural integrity and dissolution of the CeO₂ matrix, *J. Nucl. Mater.*, 2017, **484**, 332–338.
- 21 Yu. A. Teterin and A. Yu. Teterin, The structure of X-ray photoelectron spectra of light actinide compounds, *Russ. Chem. Rev.*, 2004, **73**, 541–580.
- 22 D. Shirley, High-Resolution X-ray Photoemission Spectrum of the Valence Bands of Gold, *Phys. Rev. B: Solid State*, 1972, **5**, 4709–4714.
- 23 M. Wolcyrz and L. Kepinski, Rietveld refinement of the structure of CeOCl formed in Pd/CeO₂ catalyst: Notes on the existence of a stabilized tetragonal phase of La₂O₃ in LaPdO system, *J. Solid State Chem.*, 1992, **99**, 409–413.
- 24 M. V. Ryzhkov, N. I. Medvedeva and V. A. Gubanov, Substitution of Pb by Ag as a way to obtain new high temperature superconductors: A quantum chemical point of view, *J. Phys. Chem. Solids*, 1995, **56**, 1231–1237.
- 25 A. Rosen and D. E. Ellis, Relativistic molecular calculations in the Dirac–Slater model, *J. Chem. Phys.*, 1975, **62**, 3039–3049.
- 26 D. E. Ellis and G. L. Goodman, Self-consistent Dirac–Slater calculations for molecules and embedded clusters, *Int. J. Quantum Chem.*, 1984, **25**, 185–200.
- 27 O. Gunnarsson and B. I. Lundqvist, Exchange and correlation in atoms, molecules, and solids by the spin-density-functional formalism, *Phys. Rev. B: Solid State*, 1976, **13**, 4274–4298.



- 28 P. Pyykkö and H. Toivonen, Tables of Representation and Rotation Matrices for the Relativistic Irreducible Representations of 38 Point Groups, *Acta Acad. Abo., Ser. B*, 1983, **43**, 1–50.
- 29 D. A. Varshalovich, A. N. Moskalev and V. K. Khersonskii, *Quantum Theory of Angular Momentum*, World Scientific, 1988.
- 30 A. H. Reshak, Thermoelectric properties for AA- and AB-stacking of a carbon nitride polymorph (C_3N_4), *RSC Adv.*, 2014, **4**, 63137–63142.
- 31 G. E. Davydyuk, O. Y. Khyzhun, A. H. Reshak, H. Kamarudin, G. L. Myronchuk, S. P. Danylchuk, A. O. Fedorchuk, L. V. Piskach, M. Y. Mozolyuk and O. V. Parasyuk, Photoelectrical properties and the electronic structure of $Tl_{1-x}In_{1-x}Sn_xSe_2$ ($x = 0, 0.1, 0.2, 0.25$) single crystalline alloys, *Phys. Chem. Chem. Phys.*, 2013, **15**, 6965–6972.
- 32 A. H. Reshak, Ab initio study of TaON, an active photocatalyst under visible light irradiation, *Phys. Chem. Chem. Phys.*, 2014, **16**, 10558–10565.
- 33 I. M. Band, Y. I. Kharitonov and M. B. Trzhaskovskaya, Photoionization cross sections and photoelectron angular distributions for X-ray line energies in the range 0.132–4.509 keV targets: $1 \leq Z \leq 100$, *At. Data Nucl. Data Tables*, 1979, **23**, 443–505.
- 34 M. B. Trzhaskovskaya and V. G. Yarzhemsky, Dirac-Fock photoionization parameters for HAXPES applications, *At. Data Nucl. Data Tables*, 2018, **119**, 99–174.
- 35 J. C. Slater and K. H. Johnson, Self-Consistent-Field $X\alpha$ Cluster Method for Polyatomic Molecules and Solids, *Phys. Rev. B: Solid State*, 1972, **5**, 844–853.
- 36 M. V. Rama Rao and T. Shripathi, Photoelectron spectroscopic study of X-ray induced reduction of CeO_2 , *J. Electron Spectrosc. Relat. Phenom.*, 1997, **87**, 121–126.
- 37 E. Paparazzo, G. M. Ingo and N. Zacchetti, X-ray induced reduction effects at CeO_2 surfaces: An X-ray photoelectron spectroscopy study, *J. Vac. Sci. Technol., A*, 1991, **9**, 1416–1420.
- 38 F. Le Normand, J. El Fallah, L. Hilaire, P. Légaré, A. Kotani and J. C. Parlebas, Photoemission on 3d core levels of Cerium: An experimental and theoretical investigation of the reduction of cerium dioxide, *Solid State Commun.*, 1989, **71**, 885–889.
- 39 T. Hasegawa, S. M. Shahed, Y. Sainoo, A. Beniya, N. Isomura, Y. Watanabe and T. Komeda, Epitaxial growth of $CeO_2(111)$ film on Ru(0001): scanning tunneling microscopy (STM) and X-ray photoemission spectroscopy (XPS) study, *J. Chem. Phys.*, 2014, **140**, 044711.
- 40 Yu. A. Teterin and S. G. Gagarin, Inner valence molecular orbitals and the structure of X-ray photoelectron spectra, *Russ. Chem. Rev.*, 1996, **65**, 825–847.
- 41 R. S. Mulliken, Chemical Bonding, *Annu. Rev. Phys. Chem.*, 1978, **29**, 1–31.

